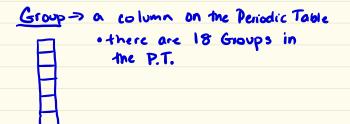
## The Periodic Table ... continued



Periods -> rows on the periodic table • there are 7 periods

What Element: Period 4, Group 17: Br Group 1, Period 1: Hydrogen Period 5, Group 9: Rh Period 3, Group 13: Al Metals

elements to the LEFT of the "stair · solid at R.Temp. (evcept Hydrogen) step" (evcept Hg/Ga)
Common Properties: (generally)

Shiny (luster)
ductile
→ they can be made into thin
sheets or really skine vire.
malleable
→ bendable

Alkali Metals - All metals in Group 1 (evcept H) - HIGHLY REACTIVE · you cannot find these elements by themselves in Nature - An Alkali Metal ( Halogen (Group17) = SALT ex. <u>NaCl</u>

## Alkaline Earth Metals

-> Metals in Group Z -> Incy are very reactive (but not as reactive as Group1) Mg, (a

Transition Metals - All metals in Group 3-12 - Very commonly known elements because they can occur in Nature Uncombined: Ag, Au, Cu, Fe, Ni, 2n

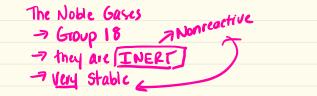
Lanthanide Series The bottom Actinide Series The bottom rows. > Man-made \* U and Pu = used for creating Nuclear energy

```
Transition Metals
    Underneath the T. Metals
   Lanthanide Series > are metals
Actinide Series > that belong
                           With the Transition Metals,
                           but they candn't fit
                         all of them.
    Nonmetal
   -they are elements to the RIGHT of
the "stair step" (and Hydrogen)
- they are usually gas or solid at
```

...

100m temperature

- Brittle (they can crumble)
- Not good conductors of heat and electricity (Metals are).
- they covalent and ionic bonds



Halogens Bonus > Group 17 J > very reactive (Flowine is the most reactive nonmetal) = Make (SALTS) with Alkali metals.

Metalloids - Iney are elements that have properties of both Nonmetals and metals: B, Si, Ge, Sb, Te, As, Po • They border the "Stair step"